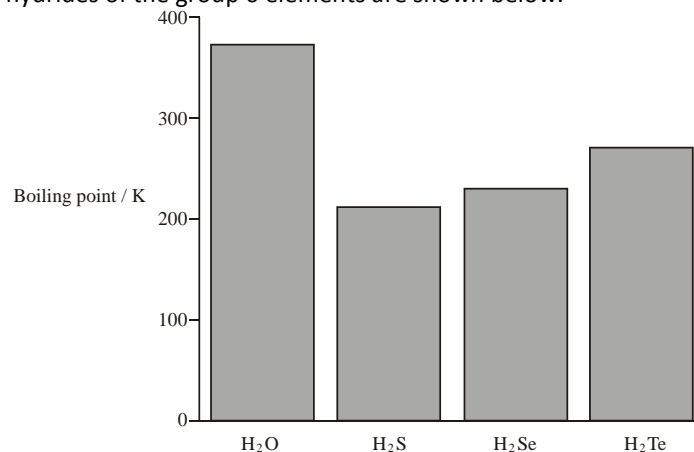


## T04D05 – 4.3 IB Practice

Name.....

1. The boiling points of the hydrides of the group 6 elements are shown below.



- (i) Explain the trend in boiling points from H<sub>2</sub>S to H<sub>2</sub>Te.

(2)

- (ii) Explain why the boiling point of water is higher than would be expected from the group trend.

(2)

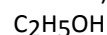
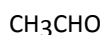
(Total 4 marks)

2. Identify the strongest type of intermolecular force in each of the following compounds.

CH<sub>3</sub>Cl .....CH<sub>4</sub> .....CH<sub>3</sub>OH .....

(Total 3 marks)

3. (i) List the following substances in order of increasing boiling point (lowest first).



.....

(2)

- (ii) State whether each compound is polar or non-polar, and explain the order of boiling points in (i).

(8)

(Total 10 marks)